

(19) World Intellectual Property  
Organization  
International Bureau



(43) International Publication Date  
8 January 2004 (08.01.2004)

PCT

(10) International Publication Number  
**WO 2004/002623 A1**

- (51) International Patent Classification<sup>7</sup>: **B01J 29/76**, 29/78, C10G 45/12, 45/34, 45/54, 45/64, 47/16, 65/12, 49/08
- (74) Agent: **GOLDBACH, Klara**; Grünecker, Kinkeldey, Stockmair & Schwanhäusser, Maximilianstrasse 58, 80538 München (DE).
- (21) International Application Number: PCT/EP2003/006585
- (81) Designated States (*national*): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NI, NO, NZ, OM, PG, PH, PL, PT, RO, RU, SC, SD, SE, SG, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, YU, ZA, ZM, ZW.
- (22) International Filing Date: 23 June 2003 (23.06.2003)
- (25) Filing Language: English
- (26) Publication Language: English
- (30) Priority Data: PA 2002 01017 28 June 2002 (28.06.2002) DK
- (84) Designated States (*regional*): ARIPO patent (GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HU, IE, IT, LU, MC, NL, PT, RO, SE, SI, SK, TR), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).
- (71) Applicant (*for all designated States except US*): **HALDOR TOPSOE A/S** [DK/DK]; Nymollevvej 55, DK-2800 Lyngby (DK).
- (72) Inventors; and
- (75) Inventors/Applicants (*for US only*): **HANSEN, Jens, Anders** [DK/DK]; Sollerødvej 90 A, DK-2840 Holte (DK). **BLOOM, Niels, Jørgen** [DK/DK]; Brunebjerg 24, DK-Hilerod (DK). **BYBERG, Birgitte, Raun** [DK/DK]; Fuglsangsvej 13, DK-3460 Birkerød (DK). **WARD, John, W.** [US/US]; 19002 Gordon Lane, Yorba Linda, CA 92686-4214 (US).
- Published:  
— with international search report  
— with amended claims
- For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.*

(54) Title: CATALYST COMPRISING ZEOLITE BETA AND ITS USE IN HYDROCARBON CONVERSION PROCESS

(57) Abstract: A hydrocarbon conversion catalyst comprising a modified beta zeolite, an amorphous inorganic oxide and a hydrogenation component wherein the said catalyst support has an NH<sub>3</sub>-AI of less than 3.5 and/or an IEC-AI of less than 3.7.

BEST AVAILABLE COPY

WO 2004/002623 A1

## CATALYST COMPRISING ZEOLITE BETA AND ITS USE IN HYDROCARBON CONVERSION PROCESS

## TECHNICAL FIELD

This invention relates to the refining of petroleum hydrocarbons into products of greater utility and higher value as compared with the feedstock.

Conversion of hydrocarbons into useful products has been practised for many years using catalytic materials. In recent years, the use of catalysts containing zeolite materials has been found in many cases to be superior to the catalysts containing amorphous inorganic oxide materials such as alumina, silica-alumina and the like. Many zeolitic materials have been found to be useful as catalysts, but depending on the specific process, zeolites such as Y, X, omega, ZSM-5, beta and L have been found to be particularly advantageous.

This invention relates in general to catalytic compositions useful for hydrocarbon conversion and in particular to their use in hydrocarbon cracking reactions such as catalytic cracking and hydrocracking. More particularly the invention relates to compositions comprising certain zeolites and metal hydrogenation components and their use in hydrocracking processes. In particular, the invention relates to catalytic compositions comprising beta zeolite and a hydrogenation component and their use in hydrocracking of hydrocarbons. The invention moreover relates to the composition and use of catalytic compositions comprising specifically modified beta zeolites, which give enhanced performance in hydrocarbon conversion reactions especially in the conversion of hydrocarbon feedstocks to middle distillates.

Often petroleum refiners produce desirable products, such as gasoline and middle distillates, by catalytically hydrocracking high boiling hydrocarbons into product hydrocarbons of lower average molecular weight and boiling point. Hydrocracking is generally accomplished by contacting in an appropriate reactor vessel a gas oil or other hydrocarbon feedstock with a suitable hydrocracking catalyst under appropriate conditions including an elevated temperature and an elevated pressure and the presence of hydrogen, such that a hydrocarbon product is obtained containing a substantial portion of a desired product boiling in a specified range, as for example a gasoline boiling in the range of 85°C to 215°C or middle distillate boiling in the range of 150°C to 425°C.

Usually, hydrocracking is practised in a single reaction vessel or several in series utilising a single catalyst. In such a scenario, the catalyst not only hydrocracks the hydrocarbon feedstock but simultaneously or sequentially converts the organonitrogen and organosulphur containing compounds into ammonia and hydrogen sulphide. Some isomerisation of normal or near normal paraffins will also take place simultaneously.

Hydrocracking can also be performed in conjunction with hydrotreating usually by a method referred to as "integral operation". In this process, the hydrocarbon feedstock, usually a gas oil containing a substantial proportion of components boiling above a desired end point as for example 215°C in the case of certain gasolines, is introduced into a catalytic hydrotreating zone wherein in the presence of a suitable catalyst such as a zeolite or sieve-free, particu-

late catalyst comprising a Group VIII metal component and a Group VIB metal component on a porous, inorganic, refractory oxide catalyst support most often composed of alumina, and under suitable conditions including an elevated temperature (e.g. 250°C to 540°C) and an elevated pressure (e.g. 0.7 to 35 MPa) and with hydrogen as a reactant, the organonitrogen components and the organosulphur components contained in the feedstock are converted to ammonia and hydrogen sulphide, respectively. Subsequently, the entire effluent removed from the hydrotreating zone is treated in a hydrocracking zone maintained under suitable conditions of elevated temperature, pressure and hydrogen partial pressure and containing a suitable hydrocracking catalyst, such that a substantial conversion of high boiling feed components to products components boiling below the desired end point is obtained. Usually, the hydrotreating and hydrocracking zones in integral operation are maintained in separate reactor vessels, but on occasion it may be advantageous to employ a single, downflow reactor vessel containing an upper bed of hydrotreating catalyst particles and lower bed of hydrocracking particles. Examples of integral operation may be found in US Patent Nos. 3,132,087; 3,159,564; 3,655,551 and 4,040,944, all of which are herein incorporated by reference in their entireties.

When two catalysts in two separate vessels are used, it is often desirable to fractionate (or separate) the products of the first reactor (hydrotreating) so as to remove the produced ammonia, hydrogen sulphide and light gaseous hydrocarbons from the feed to the hydrocracking reactor. Such separation can also be made when two similar catalysts are used.

In some integral operation refining processes, and especially those designed to produce gasoline from the heavier gas oils, a relatively high proportion of the product hydrocarbons obtained from integral operation will have a boiling point above the desired end point. For example, in the production of a gasoline product boiling in the C<sub>4</sub> to 215°C range from a gas oil boiling entirely above 300°C, it may often be the case that as much as 30 to 60 percent by volume of the products obtained from integral operation boil above 215°C. To convert these high boiling components to hydrocarbon components boiling below 215°C, the petroleum refiner separates the 215°C+ high boiling components from the other products obtained in integral operation usually after first removing ammonia by a water washing operation, a hydrogen-containing recycle gas by high pressure separation, and an H<sub>2</sub>S-containing C<sub>1</sub> to C<sub>3</sub> low BTU gas by low pressure separation. This 215°C+ boiling bottom fraction is then subjected to further hydrocracking either by recycle to the hydrocracking reactor in single stage operation or by introduction into a second hydrocracking zone wherein yet more conversion to the desired C<sub>4</sub> to 215°C product takes place.

In the foregoing two stage process, the two hydrocracking reaction zones often contain hydrocracking catalysts of the same composition. One catalyst suitable for such use is disclosed as Catalyst A in Example 16 of US Patent Nos. 3,897,327 and 3,929,672, both of which are herein incorporated by reference in their entireties, which catalyst is comprised of a palladium-exchanged, steam-stabilised Y zeolite hydrocracking component. But although the catalysts used in the two hydrocracking reaction zones may have the

same composition and the same catalytic properties, the hydrocracking conditions required in the second hydrocracking reaction zone are less severe than those required in the first. The reason for this is that ammonia is not present in the second hydrocracking reaction zone (due to water washing) whereas a significant amount of ammonia is present in the first hydrocracking zone. To account for the difference in operating conditions, it is believed that ammonia neutralises or otherwise interferes with the acidity of the zeolite in the catalyst of the first reaction zone, thereby forcing the refiner to employ relatively severe conditions for operation as for example increased temperature. On the other hand, in the ammonia-deficient atmosphere of the second hydrocracking reaction zone, high conversions to the desired product are obtainable under relatively moderate conditions often with an operating temperature about 50°C to 110°C lower than that required in the first hydrocracking reaction zone.

Further description of two-stage hydrocracking operations may be found in US Patent Nos. 4,429,053 and 4,857,169 herein incorporated by reference in their entireties, which patents provide process flow sheets for typical two-stage hydrocracking processes.

Although there exist several types of commercial hydrocracking catalysts, which can be used effectively in single stage hydrocracking or either the first, second or both stages of the above-discussed two-stage hydrocracking process, there is always a demand for new catalysts with superior overall activity, selectivity and stability for

producing gasoline and/or middle distillate via hydrocracking.

5 The general object of the present invention is directed to novel catalysts and catalyst supports and novel acid catalysed chemical conversion processes employing such catalysts to promote the desired reactions. The present invention is particularly directed to hydrocarbon conversion catalysts and hydrocarbon conversion processes employing  
10 such catalysts comprising a hydrogenation component(s) on a catalyst support comprising zeolite beta modified to a specific acidity.

#### SUMMARY OF THE INVENTION

15 The present invention provides a composition for use in the acid catalysed chemical conversion of feedstocks containing organic compounds into reaction products, particularly in the acid catalysed conversion of hydrocarbons and most especially in hydrocracking. The composition, which may be a  
20 catalyst and/or catalyst support, comprises a modified zeolite beta. One or more amorphous, inorganic refractory oxides such as alumina, silica-alumina or other inorganic oxides may also be present in the composition. For hydro-  
25 cracking purposes, the catalyst requires a hydrogenation component such as one or more Group VIB and/or Group VIII metal components, the hydrogenation components generally being dispersed on the catalyst support material composed of the zeolite beta and amorphous oxides.

30

The catalysts of the invention are characterised specifically by having a catalyst support with an  $\text{NH}_3$ -AI less than 3.5 and/or an IEC-AI less than 3.7.

5

#### DETAILED DESCRIPTION OF THE INVENTION

Zeolite beta is a crystalline zeolite whose composition and X-ray powder diffraction analysis are disclosed in US Patent No. Reissue 28,341, herein incorporated by reference in  
10 its entirety. This zeolite beta is a large pore zeolite having a pore size above 7.0 Angstroms and a Constraint Index below 2, preferably between 0.6 and 1.0. The Constraint Index of a zeolite is a convenient measure of the extent to which a zeolite provides access to molecules of varying  
15 sizes to its internal structure. Zeolites, which provide a highly restricted access to and egress from its internal structure have a high value for the Constraint Index, and zeolites of this kind usually have pores of small size. On the other hand, zeolites, which provide relatively free ac-  
20 cess to the internal zeolite structure, have a low value for the Constraint Index. The method by which Constraint Index is determined is described fully in US Patent No. 4,016,218, the disclosure of which is herein incorporated by reference in its entirety.

25

Procedures for measuring the acidity of a zeolite by  $\text{NH}_3$  TPD and ion exchange capacity have been disclosed in US Patent No. 4,894,142 to Steigleder and US Patent No. 4,612,108 to Bezman et al., all of which are herein incor-  
30 porated by reference in their entireties. Temperature-programmed desorption of ammonia ( $\text{NH}_3$  TPD) is widely used and is considered to be an effective procedure for deter-



mining the acidity strength of zeolites. Ion exchange capacity or IEC is intended to denote the number of active cation sites (acidity) in the zeolite.

5 Zeolite beta is prepared, in general, as an aluminosilicate zeolite having a  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of at least 10. The zeolite betas have a crystal size of about 0.1 to 5 micron, a surface area of about 400 to 800, a cyclohexane adsorption capacity above 15 g/100 g and a water vapour sorptive  
10 capacity at 25°C and a  $p/p_0$  value of 0.10 of greater than 5 weight percent. As used herein " $p/p_0$ " represents the water vapour partial pressure to which the zeolite beta is exposed divided by the water vapour partial pressure at 25°C.

15 Preparation of zeolite beta has been disclosed in US Patent Reissue No. 28,341. Standard preparation of zeolite beta having a  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of 25-30 have been described in the 2<sup>nd</sup> revised edition of "Verified Syntheses of Zeolitic Materials" by H. Robson (Editor) and K.P. Lillerud (XRD  
20 Patterns), Elsevier 2001, which is herein incorporated by reference in its entirety. Zeolite beta can be obtained commercially from companies like Tosoh Corporation, Japan or Zeolyst International, The Netherlands or Süd-Chemie AG, Germany.

25 As initially prepared, zeolite beta is usually in the alkali metal form and contains an organic templating agent. In this form, the zeolite has low if any catalytic activity for promoting acid catalysed conversion reactions, e.g.  
30 cracking reactions. Accordingly, the zeolite is generally converted to more active forms by calcination to decompose and drive off the templating agent followed by base ex-

change with ammonium cations to substantially reduce the alkali metal content finally followed by another calcination to convert the ammonium-exchanged zeolite to the hydrogen form. For zeolite beta initially prepared in the sodium form, the preferred sodium content upon conversion to an active form is below 1.0 percent by anhydrous weight, preferably below about 0.5 percent by anhydrous weight, calculated as  $\text{Na}_2\text{O}$ .

Publications which further discuss the properties of zeolite Beta include US Patent Nos. 3,923,641; 4,676,887; 4,812,223; 4,486,296; 4,601,993 and 4,612,108, all of which are herein incorporated by reference in their entireties.

In accordance with the invention, it has now surprisingly been found that catalysts containing a metal hydrogenation component and certain modified zeolite beta with a certain ion exchange capacity acidity index and  $\text{NH}_3$  TPD acidity index are significantly more selective for the production of middle distillate than comparable Y zeolites which for many years have and still form the component of choice in many hydrocracking catalysts.

The use of beta zeolite in hydrocracking catalysts has been reported over several years although there appears to be little if any commercial exploitation.

In US Patent No. 3,923,641 Morrison reported hydrocracking of  $\text{C}_5$  naphtha into  $\text{C}_4$  hydrocarbons using zeolite beta. Gas oil feeds are not contemplated as feedstocks and no attention is given to middle distillate. No mention of dealumination or  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio is made. Furthermore, no

mention of zeolite  $\text{NH}_3$  TPD or ion exchange capacity is made.

US Patent No. 5,128,024 to La Pierre discloses beta zeolite  
5 containing catalyst with a  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio up to  
280:1 for simultaneous hydrocracking and dewaxing. US Patent  
No. 5,980,859 to Gajda et al. discloses the modification  
of beta zeolite by steaming and extraction with ammonium  
nitrate by ion exchange. The modified zeolite is characterised  
10 by IR spectroscopy. No measurements of the  $\text{NH}_3$  TPD or IEC  
are reported. The disclosure lists uses for hydrocracking,  
catalytic cracking, isomerisation, transalkylation and  
alkylation although only its use for alkylation and transalkylation  
is illustrated. No relationship of  
15 performance and acidity is disclosed and no specific reference  
to hydrocracking is made. The use is specifically directed to  
benzene alkylation.

US Patent No. 5,972,832 to Shi et al. discloses a hydrocracking  
20 catalyst containing nickel, tungsten, fluoride, a zeolite and  
alumina. The carrier contains 10-90% of a zeolite which has an  
acidic strength value of 1.0-2.0 mmole/g determined by ammonia  
TPD. Faujasite (Y), mordenite, ZSM-5 and omega zeolite are  
specified. No acidity strength  
25 of beta zeolite is reported. No variation of the zeolite  
acidity is reported for the same family and no acidity of the  
catalyst is reported. No dependence of the hydrocracking  
activity and selectivity on acidity and ion exchange capacity  
is reported. No illustration of the utilisation of  
30 beta zeolite or its comparative performance in hydrocracking  
is illustrated. Furthermore, no disclosure of ion exchange  
capacity is given.

US Patent No. 5,284,573 to La Pierre et al. discloses the use of beta zeolite with a  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of at least 50 and even at least 250. A ratio of 500:1 may be used. A process for hydrotreating and hydrocracking is claimed. However, no disclosure of  $\text{NH}_3$  TPD or ion exchange capacity is made.

US Patent No. 4,612,108 to Angevine et al. discloses hydrocracking with beta zeolite directed to middle distillate selectivity. The disclosure illustrates an improvement by using a graded bed in which the concentration of beta zeolite in the catalyst increases with bed depth. Although  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of the beta zeolite greater than 30 is disclosed, no properties of the zeolite containing catalyst are given. No  $\text{NH}_3$  TPD or ion exchange capacity data are disclosed.

US Patent No. 4,401,556 to Bezman et al. discloses the use of Y zeolite catalysts for hydrocracking in which the zeolite has a preferred ion exchange capacity (IEC) no greater than 0.07. No impact of IEC on hydrocracking activity is reported, no  $\text{NH}_3$  TPD is reported and no disclosure of beta zeolite is made.

US Patent No. 4,894,142 to Steigleder discloses Y zeolites with a  $\text{NH}_3$  TPD of less than about 2.0 mmole/g, preferably less than 1.5 mmole/g. It states that the selectivity of the modified Y zeolite to produce middle distillate correlates to the acid strength (Col 4 and Col 5). The acid strength is measured by the  $\text{NH}_3$  TPD value. The patent teaches how to reduce the acid strength by calcination in dry air. TPD  $\text{NH}_3$  values <2.0 mmole/g and preferably <1.50

mmole/g are taught along with an IEC of  $>0.07$  defined in US Patent No. 4,401,556. No mention of beta zeolite is made.

US Patent No. 4,820,402 to Partridge et al. discloses the  
5 use of zeolite beta for hydrocracking with  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio up to 200 for producing high boiling distillate products. No reference to IEC or  $\text{NH}_3$  TPD is made. The data illustrate a continuous change in distillate selectivity with the  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio for a range of zeolites such as  
10 Y, ZSM-5, beta and ZSM-20. No differentiation between beta and Y with  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio is disclosed. In fact a smooth correlation exists for  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratios from 3 to 300 suggesting that the  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio is dominating. Large pore zeolites with  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratios of  
15 200:1 are said to give good middle distillate selectivities. No reference to IEC or  $\text{NH}_3$  TPD is made.

#### DISCLOSURE OF INVENTION

20 The present invention is specifically directed to catalysts and catalyst support comprising beta zeolites which have been modified so that the  $\text{NH}_3$  TPD acidity index ( $\text{NH}_3\text{-AI}$ ) of the catalyst support is less than 3.5 and preferably less than about 2.3. The zeolite beta has also been modified so  
25 that the IEC acidity index (IEC-AI) of the catalyst support is less than 3.7, preferably less than about 2.7. The acidity indices  $\text{NH}_3\text{-AI}$  and IEC-AI are defined in the section covering the examples.

30 A synthesised beta zeolite can be modified by any known method or other method which results in a decrease in  $\text{NH}_3$  TPD and IEC of the original zeolite to one within the scope

of the invention. Many such methods for producing modified zeolites have been reported.

Publications, which further discuss the properties of beta zeolite, include US Patent Nos. 3,923,641; 4,676,887; 4,812,223; 4,486,296; 4,601,993 and 4,612,108. Specific illustrations of beta zeolite modifications can be found in the patent literature. Illustrative treatments of beta zeolite to produce higher  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio can be found in for example US Patent No. 4,820,402 (Partridge et al.) and US Patent No. 4,430,516 (La Pierre et al.) who teach various treatments illustrated by extraction, hydrothermal treatment and ion exchange. Other methods involving direct synthesis have been described in e.g. US Patent No. 5,554,356 to Saxton et al., US Patent No. 5,164,170 to Rubin. US Patent No. 4,923,690 to Valyocsik describes a method of synthesising beta zeolite which directly produce higher  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio than those originally disclosed in US Patent No. 3,308,069 (Wadinger). All of these patent disclosures are incorporated by reference in their entireties.

Although the modified beta zeolite is the key component of this invention, the catalyst or catalyst support will generally comprise the modified zeolite intimately mixed with an acidic amorphous component and optimally a binder. The amorphous inorganic oxide can be chosen from the well known acidic oxides such as alumina, silica, titania, magnesia, zirconia, boria, phosphorous oxides along with combinations and the like. The composition of the catalyst support may contain from about 0.5 to 50 wt% modified zeolite beta

preferably 1 to 30 wt% and more preferably 1 to 15 wt% modified zeolite beta.

5 The  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of the modified (M) zeolite beta should be at least 50 and more preferably 100.

10 The modified zeolite beta and amorphous component required in the catalyst and catalyst support of the invention are embodied into particles which contain both components. Perhaps the most convenient method for physically integrating the two components into individual particulates is to comull a wetted mixture of the components and then extrude the comulled material through a die having small openings therein of desired cross-sectional size and shape, e.g.

15 circle, trilobal clover-leaf, quadrolobal clover leafs etc., breaking or cutting the extruded matter into appropriate lengths, drying the extrudates and then calcining at a temperature, e.g.  $480^\circ\text{C}$ , or higher to produce a material suitable for use in high temperature chemical conversion

20 reactions. At present it is preferred that the catalyst be produced in cylindrical form, however, as stated above other cross-sectional shapes are possible such as clover-leafs of polylobal design, for example trilobal or quadrolobal shapes, as shown for example in Figs. 8 and 10, respectively, in US Patent No. 4,028,227 herein incorporated by reference in its entirety. Typically, the amorphous oxides besides contributing to the catalytic properties of the catalyst support also serve as a binder for the modified zeolite beta. Alumina and other conventional amorphous, inorganic refractory oxide binder components may be

30 desired.

Regardless of whether an amorphous, inorganic refractory oxide component is used as a binder material to hold the zeolite beta and amorphous oxides together in the catalyst support, it will be understood that other such components  
5 can also be incorporated into the comulled mixture including for example inorganic refractory oxide diluents, which may or may not possess some type of catalyst activity. Examples of such diluents include clays, alumina, silica-alumina and a heterogeneous dispersion of finely divided  
10 silica-alumina particles in an alumina matrix, which dispersion is described in detail in US Patent Nos. 4,097,365; 4,419,271 and 4,857,171, the disclosures of which are herein incorporated by reference in their entireties. Additionally and alternatively, hydrogenation component precursors  
15 can also be comulled into the mixture, as will be discussed in more detail hereinafter.

It will be further understood that producing the catalyst and catalyst support of the invention in extrudate form,  
20 while certainly the most highly preferred method is still but one option available to those skilled in the art. The catalyst and its support may also be produced in tablet, granules, spheres, and pellets as desired, by any known method for combining zeolites with a porous, inorganic refractory oxide component.  
25

The catalyst of the invention can be used for converting hydrocarbons and other organic compounds into more valuable reaction products by acid catalysed reactions such as alkylation, transalkylation, dealkylation, isomerisation, dehydrocyclisation, dehydrogenation, hydrogenation, cracking, hydrocracking, dewaxing, hydrodewaxing, oligomerisation,  
30



aromatisation, alcohol conversion reactions, the conversion of syngas into mixtures of hydrocarbons and the like. When the catalyst or catalyst support contains modified zeolite beta but no hydrogenation components, it is useful for any of a number of acid-catalysed hydrocarbon conversion reactions in which hydrogen is not an added reactant, e.g. isomerisation, alkylation, transalkylation, cracking, dewaxing, oligomerisation etc. However, since the main benefit of the invention as presently contemplated is in hydroprocessing such as hydrocracking, a process in which hydrogen is an added reactant, the catalyst for this purpose will further require one or more hydrogenation components, in which case the portion of the catalyst exclusive of any hydrogenation metal components is considered the catalyst support upon which the hydrogenation component(s) is dispersed.

Whether the modified zeolite beta particulates, which may optionally include an inorganic refractory oxide binder and/or diluent, are used as the catalyst itself or as the catalyst support (or support component) for hydrogenation metals, the amounts of the modified zeolite beta and other components present will normally depend on the particular process in which the particles are used.

When the modified zeolite beta particulates are used for selectively producing middle distillates in a hydrocracking process, the catalyst support will typically on a dry basis contain less than 50 wt%, preferably less than 30 wt% and more preferably less than 15 wt% modified zeolite beta with at least 50 wt%, preferably 100 wt% of the balance being an

amorphous, inorganic refractory oxide binder and diluent in combination.

For use in hydroprocessing, such as hydrocracking, the catalyst contains one or more hydrogenation components containing metals selected from the group consisting of Group VIB and/or Group VIII of the periodic table of elements, such components typically being in the form of the free metals or their respective oxides and sulphides the latter two being most preferred. As used herein "Periodic Table of Elements" refers to the version found in the inside front cover of the "Handbook of Chemistry and Physics", 59<sup>th</sup> Edition, published in 1979 by the Chemical Rubber Company. The platinum group (or noble) metals of the Group VIII metals may be used, but preference is at present for the base (or non-noble) metals, e.g. nickel and cobalt in particular, and nickel most preferably of all. Of the group VIB metals, molybdenum and tungsten are preferred with molybdenum being most preferred when the catalyst is to be used in gasoline hydrocracking and tungsten being most preferred when the catalyst is to be used in middle distillate hydrocracking. The most highly preferred catalyst contains both a non-noble Group VIII metal component and a Group VIB metal component, most preferably nickel and molybdenum or nickel and tungsten in combination.

The hydrocracking catalysts of the invention contain at least 0.2 wt% of the hydrogenation components calculated as the metals. If noble metals are used, the hydrogenation components are generally present in a relatively low proportion, e.g. 0.2 to 2 wt%. For the base or non-noble metals, the proportions are generally higher. Non-noble Group

VIII metal components are typically employed in proportions between about 2 and 15 wt%, preferably between 3 and 10 wt%, calculated as the respective monoxide. The Group VIB metal component is generally employed in proportions of about 5 to 35 wt%, preferably in the range of 8 to 30 wt%, calculated as the respective trioxide. It is to be understood that the proportions given above for the hydrogenation metal components are based on the finished catalyst whereas the proportions expressed above for the modified zeolite beta particulates are values in the absence of the hydrogenation metal component, i.e. for the catalyst support only. For purposes herein, the term "catalyst support" is defined as all materials in the catalyst except the hydrogenation metal components.

The hydrogenation components may be incorporated into the catalyst in any of many ways known in the art of combining hydrogenation components with catalyst supports containing zeolites. One such method is to first prepare the catalyst support, for example, as an extrudate containing zeolite beta and an amorphous inorganic refractory oxide in calcined form and then impregnating the catalyst support with solutions containing the desired metal(s) in dissolved form. Calcination in air, typically in the absence of added steam at an elevated temperature, e.g. above 425°C, preferably above 475°C, produces the desired catalyst containing metals in oxide form. Likewise, in another embodiment, the desired metal(s) are introduced by comulling a compound containing such metal(s) in the zeolite beta amorphous oxide mixture previously described followed by shaping (e.g. extrusion through a die), drying and calcining in the substantial absence of steam, e.g. at a temperature between

about 425° and 550°C to produce the oxide form of the catalyst. For one preferred catalyst, the comulling is effected with ammonium heptamolybdate as the source of molybdenum and nickel nitrate as the source of nickel with both compounds generally being introduced into the mulling mixture in the form of an aqueous solution. Other metals can be similarly introduced in dissolved aqueous form likewise, non-metallic elements, e.g. phosphorus.

Catalysts, which contain hydrogenation components in the oxide form as described above, are generally treated to convert the metals to the sulphide form prior to the use of the catalysts in hydrocracking. This can be accomplished by presulphiding the catalyst prior to use at an elevated temperature, e.g. 150 to 375°C with for example a mixture consisting of 10 volume percent H<sub>2</sub>S and 90 volume percent H<sub>2</sub>. Alternatively, the catalyst can be presulphided ex situ by various sulphiding processes; as an illustration, see "Sulphicat<sup>R</sup>: Offsite Presulphiding of Hydroprocessing Catalysts from Eurocat" by J.H. Wilson and G. Berrebi, Catalysts 87, Studies in Surface Science and Catalysts #38 page 393. Alternatively, the sulphiding is accomplished in situ, i.e. by using the catalyst in the oxide form to hydrocrack a hydrocarbon feedstock containing sulphur compounds under hydrocracking conditions including elevated temperature and pressure and the presence of hydrogen.

Hydrocracking catalysts in accordance with the invention are useful in the conversion of a wide variety of hydrocarbon feedstocks to a hydrocarbon product of lower average boiling point and/or molecular weight. The feedstocks that may be subjected to hydrocracking by the methods of the in-

vention include all mineral oils and synthetic oils (e.g. shale oil, tar sand products etc.) and fractions thereof. Illustrative feedstocks include straight run gas oils, vacuum gas oils, coker gas oils and catcracker distillates.

5 The typical hydrocracking feedstock, however, contains a substantial proportion of components usually at least 50 percent by volume, often at least 75% by volume boiling above the desired end point of the products, which end point in the case of gasoline, will generally be in the

10 range of about 190°C to 215°C and in the case of middle distillates will typically be in the range of 340°C to 425°C. Usually, the feedstock will also contain gas oil components boiling above 285°C with highly useful results being achievable with feeds containing at least 30 percent

15 by volume of components boiling between 315°C and 600°C.

For best results in hydrocracking, the catalyst of the invention will be employed as a fixed bed of catalytic particulates in a hydrocracking reactor vessel into which hydrogen and the feedstock are introduced and passed in a

20 downwardly direction. Operating conditions in the reactor vessel are chosen so as to convert the feedstock into the desired product, which in one preferred embodiment is a hydrocarbon product containing a substantial proportion of

25 gasoline components boiling, for example in the 85°C to 215°C range. However, other products such as middle distillates boiling in the 150°C to 425°C range may also be highly desired, and conditions must be adjusted according to the product (or distribution of products) desired. The

30 unconverted oil, e.g. hydrocarbons boiling at a temperature higher than that of the middle distillate fraction may be used for production of lube oil blending stocks, fluid

cracker feedstocks or ethylene cracker feedstock. The exact conditions required in a given situation will depend upon the nature of the feedstock, the particular catalyst composition utilised, and the desired product(s). In general, the conditions of operation for hydrocracking will fall into the following usual and preferred ranges.

TABLE 1

Conditions	Usual	Pre-ferred
Temperature, °C	230-455	260-430
Pressure, Mpa	4-30	5-20
LHSV, h <sup>-1</sup>	0.3-5.0	0.5-3.5
H <sub>2</sub> /feedstock, Nm <sup>3</sup> /m <sup>3</sup>	200-2000	500-1500

10

The foregoing Table 1 shows the suitable and preferred hydrocracking conditions for single stage or for each stage of a two stage operation. It will be understood, however, that the operating conditions in the two stages of the two stage process are not necessarily identical. In fact, as mentioned before, the primary difference in conditions in the two hydrocracking reactor vessels of two stage operation is the presence of substantial amounts of ammonia often greater than about 2000 vol ppm in the first stage, and its essential absence, i.e. less than 200 vol ppm and preferably less than about 20 vol ppm in the second, allowing for less severe conditions in the second stage. There may, however, be yet other differences in conditions in any particular situation.

20

25

Based on the presently available data, the catalysts of the invention, when compared with Y zeolite catalysts of similar NH<sub>3</sub>-AI and IEC-AI, are found to be substantially more

selective for middle distillate production. These achievements and others are proven in the following examples, which are provided for illustration purposes and not to limit the invention as defined by the claims.

5

#### EXAMPLES

The key parameter in defining the zeolites are measured by the following techniques:

10

##### AMMONIA TEMPERATURE PROGRAMMED DESORPTION (NH<sub>3</sub> TPD)

A 0.1 g sample of zeolite is heated in helium at 500°C for 2 hours. The pretreated zeolite is then cooled to 150°C. The sample is then saturated at this temperature by a mixture of 2 vol% ammonia in helium for ¼ hour. The saturated zeolite is then purged with helium for 2 to 3 hours. Immediately after this period, the exit gas is directed to a bubble flask containing 100 ml of a 0.65 N boric acid solution and 100 ml water in order to collect the desorbed ammonia, while the temperature is raised to 850°C with a temperature ramp of 10°C/min. The boric acid solution is then titrated with a 0.01 N HCl solution until the pH equals 4,90. At this point the amounts of HCl added equals the amount of NH<sub>3</sub> desorbed from the zeolite sample, and the acidity is reported as mmole NH<sub>3</sub> desorbed per gram of sample after the determination.

20

25

##### ION EXCHANGE CAPACITY (IEC)

10 grams of zeolite are contacted three times at 25°C for a period of one hour each time with a fresh aqueous ion exchange solution containing 10 g of NaCl per liter of solution (100 ml of solution is used for each gram of zeolite).

30

After this contact, the zeolite is contacted for one hour with 100 ml hot pure water for each gram of zeolite (60-70°C) in order to remove excess of sodium. The sodium content after calcination at 600°C is then determined and reported as the ion exchange capacity.

#### NH<sub>3</sub> TPD ACIDITY INDEX (NH<sub>3</sub>-AI)

The NH<sub>3</sub>-AI is defined as the amount of mole NH<sub>3</sub>/g zeolite (measured by the NH<sub>3</sub> TPD method) of the zeolite used in the catalyst support multiplied by the weight percentage of the zeolite in the catalyst support normalised by a catalyst support containing 1 wt% of a standard zeolite beta (CP 814E from Zeolyst International) having a specified SiO<sub>2</sub>:Al<sub>2</sub>O<sub>3</sub> molar ratio of 25 and a NH<sub>3</sub>-TPD of 0.90 mmole NH<sub>3</sub>/g zeolite (all on dry weight basis).

$$\text{NH}_3\text{-AI} = \frac{(\text{mole NH}_3/\text{g zeolite}) * (\text{zeolite content in catalyst support, wt\%})}{(\text{mmol NH}_3/\text{g standard zeolite beta}) * (1 \text{ wt\%})}$$

#### ION EXCHANGE CAPACITY ACIDITY INDEX (IEC-AI)

The IEC-AI is defined as the amount of wtppm of sodium in the zeolite (measured by the IEC method) used in a catalyst support multiplied by the weight percentage of zeolite in the catalyst support normalised by a catalyst support containing 1% of a standard zeolite beta (CP 814E from Zeolyst International) having a specified SiO<sub>2</sub>:Al<sub>2</sub>O<sub>3</sub> molar ratio of 25 and an IEC of 18,500 wtppm sodium (all on dry weight basis).



$$\text{IEC AI} = \frac{\text{wtppm sodium}) * (\text{zeolite content in catalyst support, wt\%})}{(\text{wtppm sodium in standard zeolite beta}) * (1 \text{ wt\%})}$$

5

**Example 1**

A catalyst was prepared by mixing 10 wt% of a commercially available zeolite beta having a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of about 1500, a  $\text{NH}_3$  TPD of 0.077 mmole  $\text{NH}_3/\text{g}$  zeolite and an ion exchanged capacity (IEC) of 630 wtppm sodium with 35 wt% amorphous silica-alumina (high alumina type) and 55 wt% gamma alumina. The resulting mixture was extruded to form 1/16" extrudates. The extrudates were dried and calcined at 550°C for 2 hours in air. The calcined extrudates were impregnated with an aqueous solution containing nickel nitrate and ammonium metatungstate. The wet extrudates were dried and calcined at 550°C for 2 hours in air. The final catalyst contained 6 wt% nickel oxide ( $\text{NiO}$ ) and 22 wt% tungsten oxide ( $\text{WO}_3$ ) (all on dry weight basis).

**Examples 2-4**

The three catalysts of Examples 2-4 were prepared in the same way as in Example 1 except that other beta zeolites were used. In Example 2, a commercially available zeolite beta having a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of about 500, a  $\text{NH}_3$  TPD of 0.092 mmole  $\text{NH}_3/\text{g}$  zeolite and an IEC of 2000 wtppm sodium was used. In Example 3, a commercially available zeolite beta with a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of about 300, a  $\text{NH}_3$  TPD of 0.151 mmole  $\text{NH}_3/\text{g}$  zeolite and an IEC of 2870 wtppm sodium was used. In Example 4, a commercially available zeolite beta with a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$

molar ratio of about 100, a  $\text{NH}_3$  TPD of 0.236 mmole  $\text{NH}_3/\text{g}$  zeolite and an IEC of 5630 wtppm sodium was used.

#### Example 5

5 A sample of a commercially available zeolite Y type was steamed in 100% steam for two hours at 750°C resulting in an  $\text{NH}_3$  TPD of 0.089 mmole  $\text{NH}_3/\text{g}$  zeolite and an IEC of 1630 wtppm sodium. A portion of this modified zeolite Y was mixed with 74 wt% amorphous silica-alumina (high alumina  
10 type), and 20 wt% gamma alumina to obtain a mixture containing 6 wt% of the modified zeolite Y. The mixture was extruded, calcined and impregnated as in Example 1.

#### Example 6

15 A catalyst was prepared in the same way as described in Example 5 except that a commercially available zeolite Y type with a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of about 15, a  $\text{NH}_3$  TPD of 0.118 mmole  $\text{NH}_3/\text{g}$  zeolite and an IEC of 2105 wtppm sodium was used.

20

#### Examples 7-8

Two catalysts were prepared in the same way as described in Example 1 expect that the beta zeolite component was replaced by zeolite Y. In Example 7, a commercially available  
25 zeolite Y type with a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of about 55, a  $\text{NH}_3$  TPD of 0.229 mmole  $\text{NH}_3/\text{g}$  zeolite and an IEC of 5990 wtppm sodium was used. In Example 8, a commercially available zeolite Y type with a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of about 30, and an  $\text{NH}_3$  TPD of 0.428 mmole  $\text{NH}_3/\text{g}$  zeo-  
30 lite and an IEC of 9570 wtppm sodium was used.

The properties of the catalyst support with the calculated  $\text{NH}_3$ -AI and IEC-AI of Examples 1-8 are given in Table 2. The zeolite used in the examples have been obtained from Tosoh Corporation Zeolyst International.

5

#### Example 9

The catalysts prepared in Examples 1-8 were tested according to the following conditions. Prior to testing, the catalyst was sulphided in the presence of hydrogen and hydrogen sulphide at a temperature from 150°C to 360°C. Feed A (properties are given in Table 3) was passed on a once-through basis through an isothermal reactor containing catalyst particles uniformly mixed with carborundum. The catalyst was tested under the following operating conditions: LHSV of 0.5-1.0  $\text{h}^{-1}$ , total pressure 14.2 MPa, and a hydrogen gas/feed ratio of 1500-1800  $\text{Nl/l}$ . The temperature of the reactor was adjusted to provide 75 wt% conversion of the >360°C fraction in the feedstock. The temperature required to obtain the desired conversion level was from 380-400°C. The selectivity to middle distillate products was measured as the percentage fraction of the 360°C minus product fraction boiling in the range from 160-360°C. The conversion and selectivity were calculated from gas chromatograph boiling range analyses of the product according to ASTM D 2887.

25

The results of the testing of all catalysts are given in Table 4. The data in Table 2 and Table 4 and Figs. 1-2 show a comparison of catalysts containing modified zeolite beta versus catalysts containing zeolite Y. The catalysts of the invention are very superior in terms of selectivity to middle distillates. The data also reveal that the catalysts

30

had excellent denitrogenation and desulphurisation activities. The pour point of the >360°C product cut as compared with the pour point of the same fraction in the feedstock was reduced significantly during the processing.

5

**Example 10**

A catalyst was prepared in the same way as in Example 1 containing 10 wt% of a commercially available zeolite Y type having a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of about 15, an  $\text{NH}_3$  TPD of 0.118 mmole  $\text{NH}_3/\text{g}$  zeolite and an IEC of 2105 wtpm sodium resulting in an  $\text{NH}_3$ -AI of 1.311 and a IEC-AI of 1.249.

10

**Example 11**

A commercially available zeolite beta with a specified  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of about 300, an  $\text{NH}_3$  TPD of 0.151 mmole  $\text{NH}_3/\text{g}$  zeolite and an IEC of 2870 wtpm sodium resulting in an  $\text{NH}_3$ -AI of 1.174 and an IEC-AI of 1.086 was mixed with 38 wt% amorphous silica-alumina (high alumina type) and 55 wt% gamma alumina to obtain 7 wt% zeolite in the catalyst support. The mixture was extruded to form 1/16" extrudates. The extrudates were calcined and impregnated as in Example 1. The zeolite used in the examples have been obtained from Tosoh Corporation Zeolyst International.

15

20

25

**Example 12**

The catalysts prepared in Examples 10 and 11 were tested according to the procedure given in Example 9 except that Feed B with the properties given in Table 3 was used. The zeolite used in the examples have been obtained from Tosoh Corporation Zeolyst International.

30

The test of the catalyst containing zeolite Y prepared in Example 10 with an  $\text{NH}_3$ -AI of 1.311 and an IEC-AI of 1.249 results in a selectivity to middle distillate of 69.1%. The catalyst prepared in Example 11 containing zeolite beta  
5 with an  $\text{NH}_3$ -AI of 1.174 and an IEC-AI of 1.086 results in a selectivity to middle distillate of 71.9%.

Example 12 illustrates that the catalyst containing zeolite beta in accordance with the present invention is clearly  
10 more selective to middle distillate on a second feed than the catalyst based on zeolite Y with similar  $\text{NH}_3$ -AI and IEC-AI.

TABLE 2

15

ZEOLITE AND CATALYST PROPERTIES			
Exam- ple	Zeolite Type	$\text{NH}_3$ -AI wt%	IEC-AI wt%
1	Beta	0.856	0.341
2	Beta	1.022	1.081
3	Beta	1.678	1.551
4	Beta	2.622	3.043
5	Y	0.593	0.529
6	Y	0.787	0.683
7	Y	2.544	3.243
8	Y	4.759	5.173

**TABLE 3**  
**FEED PROPERTIES**

	FEED A	FEED B
Specific gravity	0.8785	0.8816
Sulphur, D-4294, wt%	0.84	0.55
Nitrogen, D-4629, wtppm	641	601
Hydrogen, D-4804, wt%	12.95	13.05
Aromatics, IP-391, wt%	12.9	15.4
Mono	5.1	4.6
Di	6.6	6.1
Tri+	30	30
Pour Point, D-5949, °C	13.79	11.5
Viscosity @40°C, cSt	3.229	3.506
Viscosity @50°C, cSt		
Viscosity @100°C, cSt		
<u>Distillation, °C, D-</u>		
2887		
IBP/5	168/26	209/274
10/20	4	308/344
30/40	296/32	366/383
50/60	6	399/413
70/80	354/37	425/437
90/95	1	451/462
FBP	387/40	489
	2	
>360°C, wt%	416/43	72.6
	2	
<u>&gt;360°C cut</u>	456/47	
Pour Point, D-5949, °C	7	
Specific gravity	514	
	66.7	
	36	
	0.8869	

TABLE 4

## TEST RESULTS

Exam- ple	Feed	Selectivity to middle distillates		>360°C Pour Point	Product Ni- trogen	Product Sulphur
		160-360°C				
1	A	73.9%		0°C	<1 wtppm	1 wtppm
2	A	73.5%		-1°C	<1 wtppm	2 wtppm
3	A	69.2%		7°C	<1 wtppm	4 wtppm
4	A	65.2%		7°C	<1 wtppm	3 wtppm
5	A	69.4%				
6	A	69.4%				
7	A	66.6%				
8	A	62.5%				

Although it should be more than clear from the previous how  
 5 the invention can be utilised in the industry, it can be  
 summarised as follows: The invention will find its use in  
 the petroleum refining industry, and more specifically in  
 those processes employed in the petroleum refining industry  
 that require catalysts. Catalysts prepared in accordance  
 10 with the invention are most usefully employed without hy-  
 drogenation components in such processes as catalytic  
 cracking, catalytic isomerisation, catalytic alkylation and  
 catalytic transalkylation. The catalysts of the invention

are most usefully employed with one or more hydrogenation components when catalytic hydrocracking, catalytic hydrodenitrogenation or catalytic hydrodesulphurisation is required. The best contemplated use of the catalyst of the invention is for hydrocracking of gas oils and the like to produce midbarrel products boiling in the 150°C to 420°C range, and for such hydrocracking the most preferred catalyst comprises sulphided nickel and tungsten components on a catalyst support consisting essentially of a modified beta zeolite, silica-alumina, alumina and binder.

While particular embodiments of the invention have been described, it goes without saying that the invention is not limited thereto since many obvious modifications can be made, and it is intended to include within this definition any such modifications as will fall within the scope of the appended claims.



**CLAIMS**

1. A hydrocarbon conversion catalyst comprising a modified beta zeolite, an amorphous inorganic oxide and a hydrogenation component wherein the said catalyst support has an  $\text{NH}_3\text{-AI}$  of less than 3.5, preferably less than about 2.3, and/or an  $\text{IEC-AI}$  of less than 3.7, preferably less than about 2.7.
2. A hydrocarbon conversion catalyst of claim 1, in which the modified beta zeolite has a  $\text{SiO}_2\text{:Al}_2\text{O}_3$  molar ratio of at least 50.
3. A hydrocarbon conversion catalyst of claim 1, in which the modified beta zeolite has a  $\text{SiO}_2\text{:Al}_2\text{O}_3$  molar ratio of at least 100.
4. A hydrocarbon conversion catalyst of claim 1, in which the said hydrogenation components are selected from the elements of Group VIII and/or Group VI B.
5. A hydrocarbon conversion catalyst of claim 1, in which the amorphous inorganic oxide is selected from alumina, silica, titania, zirconia, magnesia, boria, phosphorous oxides and their combinations.
6. A hydrocarbon conversion process that comprises contacting a hydrocarbon feedstock in the presence of hydrogen under hydrocarbon conversion conditions with a catalyst according to any of the proceeding claims.

**AMENDED CLAIMS**

[received by the International Bureau on 05 December 2003 (05.12.03);  
original claims 1-6 replaced by new claims 1-8 (2 pages)]

1. A hydrocarbon conversion catalyst comprising a modified beta zeolite, an amorphous inorganic oxide and a hydrogenation component, wherein the said catalyst support has an  $\text{NH}_3$ -TPD Acidity Index of less than 3.5 and/or an Ion Exchange Capacity-Acidity Index of less than 3.7.
2. A hydrocarbon conversion catalyst of claim 1, in which the  $\text{NH}_3$ -TPD Acidity Index is less than 2.3 and/or the Ion Exchange Capacity-Acidity Index is less than 2.7.
3. A hydrocarbon conversion catalyst of claim 1, in which the modified beta zeolite has a  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of at least 50.
4. A hydrocarbon conversion catalyst of claim 1, in which the modified beta zeolite has a  $\text{SiO}_2:\text{Al}_2\text{O}_3$  molar ratio of at least 100.
5. A hydrocarbon conversion catalyst of claim 1, in which the said hydrogenation components are selected from the elements of Group VIII and/or Group VI B.
6. A hydrocarbon conversion catalyst of claim 1, in which the amorphous inorganic oxide is selected from alumina, silica, titania, zirconia, magnesia, boria, phosphorous oxides and their combinations.

7. A hydrocarbon conversion process that comprises contacting a hydrocarbon feedstock in the presence of hydrogen under hydrocarbon conversion conditions with a catalyst according to any of the proceeding claims.
8. A hydrocarbon conversion process according to claim 7, in which the process comprises selective conversion of hydrocarbons to middle distillates in a hydrocracking process.

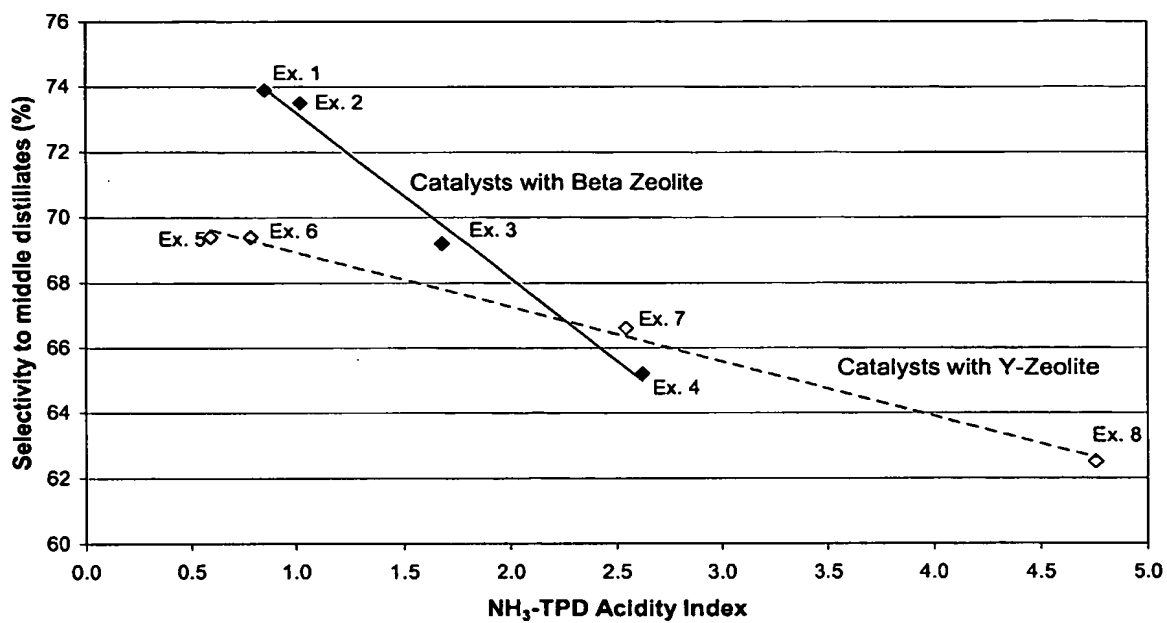
Selectivity vs.  $\text{NH}_3$ -AI

Fig. 1

## Selectivity vs. IEC-AI

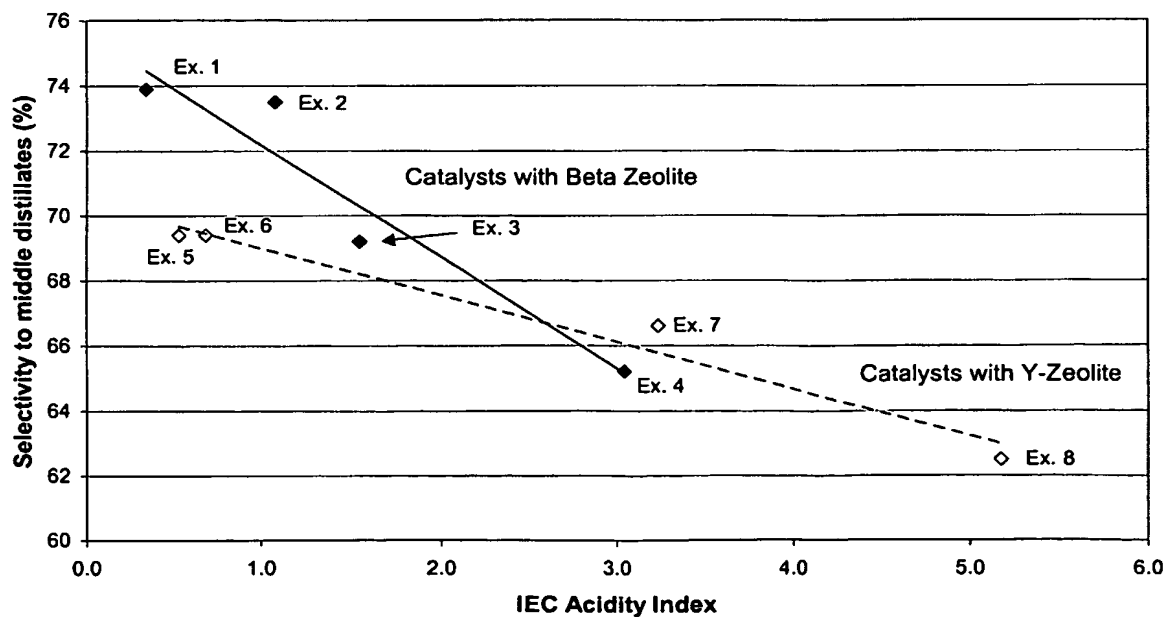


Fig. 2

# INTERNATIONAL SEARCH REPORT

International Application No  
PCT/EP 06585

## A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 B01J29/76 B01J29/78 C10G45/12 C10G45/34 C10G45/54  
C10G45/64 C10G47/16 C10G65/12 C10G49/08

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 B01J C10G

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal, WPI Data, COMPENDEX, INSPEC

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 2002/074263 A1 (KOEGLER JOHANNES HENDRIK ET AL) 20 June 2002 (2002-06-20) paragraphs '0007!', '0031!'-'0034! claims 1-3,6-8,26-31; examples & WO 03 045548 A (ABB LUMMUS GLOBAL INC.) 05 June 2003 (2003-06-05)	1-6
X	WO 01 17901 A (JANSEN JACOBUS CORNELIS ;MASCHMEYER THOMAS (NL); SHAN ZHIPING (NL)) 15 March 2001 (2001-03-15) page 13, line 8 - line 13 claims; example 8	1-5

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

### \* Special categories of cited documents :

- \*A\* document defining the general state of the art which is not considered to be of particular relevance
- \*E\* earlier document but published on or after the international filing date
- \*L\* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- \*O\* document referring to an oral disclosure, use, exhibition or other means
- \*P\* document published prior to the international filing date but later than the priority date claimed

- \*T\* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- \*X\* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- \*Y\* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- \*&\* document member of the same patent family

Date of the actual completion of the international search

26 September 2003

Date of mailing of the international search report

06/10/2003

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2  
NL - 2280 HV Rijswijk  
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,  
Fax: (+31-70) 340-3016

Authorized officer

Gosselin, D

## INTERNATIONAL SEARCH REPORT

 Internatio      Application No  
 PCT/EP      06585

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP 1 106 592 A (CHINA PETRO CHEMICAL CORP) 13 June 2001 (2001-06-13) paragraphs '0026!-'0028!, '0035!, '0037!, '0039! tables 1,2 claims 1,4-10,13-22,25 ---	1-6
X	US 5 800 698 A (TORREALBA MARIANA ET AL) 1 September 1998 (1998-09-01) column 1, line 14 - line 17 examples 1,5,9,10 claims 1,5,6,10,13,16,17,23-25,34,40,43,45,57,58 ---	1-6
X	WO 96 13563 A (MOBIL OIL CORP) 9 May 1996 (1996-05-09) claims; examples ---	1-6
X	US 4 612 108 A (ANGEVINE PHILIP J ET AL) 16 September 1986 (1986-09-16) cited in the application column 7, line 64 -column 8, line 3 claims ---	1-6
X	US 4 812 223 A (HICKEY JR KARLTON J ET AL) 14 March 1989 (1989-03-14) cited in the application column 2, line 12 - line 37 claims; examples ---	1-6
X	US 4 486 296 A (OLECK STEPHEN M ET AL) 4 December 1984 (1984-12-04) cited in the application column 6, line 25 -column 7, line 10 claims; examples ---	1-6
X	US 4 601 993 A (YEN JEFFREY H ET AL) 22 July 1986 (1986-07-22) cited in the application column 3, line 47 - line 68 examples 7-29,34-38,46-40 claims ---	1-6
X	US 5 011 593 A (WARE ROBERT A ET AL) 30 April 1991 (1991-04-30) cited in the application column 4, line 61 -column 5, line 61 claims; examples -----	1-6

FURTHER INFORMATION CONTINUED FROM PCT/ISA/ 210

## Continuation of Box I.2

The catalyst according to claims 1 to 5 is defined by reference to acidity indexes, which are defined in the application. The definitions of the acidity indexes given in the application may be regarded as clear and concise. However, no meaningful comparison with the prior art can be made, since the catalysts known in the art for the same purpose are not commonly defined by reference to acidity indexes as defined in the application. Consequently, the subject-matter of claims 1 to 6 does not meet the requirements of Article 6 PCT.

The lack of clarity is such as to render a meaningful complete search impossible. Consequently, the search has been restricted to the compositions of claims 1 to 5 and their use in a process according to claim 6 the acidity indexes having been omitted.

The applicant's attention is drawn to the fact that claims, or parts of claims, relating to inventions in respect of which no international search report has been established need not be the subject of an international preliminary examination (Rule 66.1(e) PCT). The applicant is advised that the EPO policy when acting as an International Preliminary Examining Authority is normally not to carry out a preliminary examination on matter which has not been searched. This is the case irrespective of whether or not the claims are amended following receipt of the search report or during any Chapter II procedure.

# INTERNATIONAL SEARCH REPORT

Intern application No.  
PCT/EP 03/06585

## Box I Observations where certain claims were found unsearchable (Continuation of item 1 of first sheet)

This International Search Report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. ☐ Claims Nos.:  
because they relate to subject matter not required to be searched by this Authority, namely:
  
2. ☒ Claims Nos.:  
because they relate to parts of the International Application that do not comply with the prescribed requirements to such an extent that no meaningful International Search can be carried out, specifically:  
see FURTHER INFORMATION sheet PCT/ISA/210
  
3. ☐ Claims Nos.:  
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

## Box II Observations where unity of invention is lacking (Continuation of item 2 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

1. ☐ As all required additional search fees were timely paid by the applicant, this International Search Report covers all searchable claims.
  
2. ☐ As all searchable claims could be searched without effort justifying an additional fee, this Authority did not invite payment of any additional fee.
  
3. ☐ As only some of the required additional search fees were timely paid by the applicant, this International Search Report covers only those claims for which fees were paid, specifically claims Nos.:
  
4. ☐ No required additional search fees were timely paid by the applicant. Consequently, this International Search Report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

- ☐ The additional search fees were accompanied by the applicant's protest.
- ☐ No protest accompanied the payment of additional search fees.



# INTERNATIONAL SEARCH REPORT

Information on patent family members

International

Application No

PCT/EP

06585

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
US 2002074263	A1	20-06-2002	WO 03045548 A1	05-06-2003
			US 2003017943 A1	23-01-2003
			AU 3893900 A	10-04-2001
			BR 0013787 A	27-08-2002
			CA 2384090 A1	15-03-2001
			CN 1387496 T	25-12-2002
			EP 1214272 A1	19-06-2002
			JP 2003508333 T	04-03-2003
			WO 0117901 A1	15-03-2001
WO 0117901	A	15-03-2001	AU 3893900 A	10-04-2001
			BR 0013787 A	27-08-2002
			CA 2384090 A1	15-03-2001
			CN 1387496 T	25-12-2002
			EP 1214272 A1	19-06-2002
			JP 2003508333 T	04-03-2003
			WO 0117901 A1	15-03-2001
			US 2003017943 A1	23-01-2003
			US 2002074263 A1	20-06-2002
EP 1106592	A	13-06-2001	CN 1287884 A	21-03-2001
			EP 1106592 A1	13-06-2001
			WO 0100547 A1	04-01-2001
US 5800698	A	01-09-1998	US 5612273 A	18-03-1997
			US 5817595 A	06-10-1998
			US 5916433 A	29-06-1999
WO 9613563	A	09-05-1996	AU 698961 B2	12-11-1998
			AU 3763395 A	23-05-1996
			CA 2200945 A1	09-05-1996
			EP 0788533 A1	13-08-1997
			JP 10508058 T	04-08-1998
			WO 9613563 A1	09-05-1996
US 4612108	A	16-09-1986	CA 1272152 A1	31-07-1990
			DE 3676265 D1	31-01-1991
			EP 0216443 A1	01-04-1987
			JP 62034987 A	14-02-1987
			SG 11092 G	16-04-1992
US 4812223	A	14-03-1989	NONE	
US 4486296	A	04-12-1984	AU 573630 B2	16-06-1988
			AU 3414084 A	18-04-1985
			CA 1226268 A1	01-09-1987
			DE 3463002 D1	14-05-1987
			EP 0140608 A1	08-05-1985
			JP 1785615 C	31-08-1993
			JP 4072579 B	18-11-1992
			JP 60097051 A	30-05-1985
			KR 9108565 B1	19-10-1991
US 4601993	A	22-07-1986	CA 1243995 A1	01-11-1988
			DE 3586296 D1	13-08-1992
			DE 3586296 T2	04-03-1993
			EP 0163449 A2	04-12-1985
			JP 1886448 C	22-11-1994

# INTERNATIONAL SEARCH REPORT

Information on patent family members

International

Application No

PCT/EP

88/06585

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
US 4601993	A	JP 6007926 B JP 60261544 A	02-02-1994 24-12-1985
US 5011593	A	30-04-1991	NONE

**This Page is Inserted by IFW Indexing and Scanning  
Operations and is not part of the Official Record**

**BEST AVAILABLE IMAGES**

Defective images within this document are accurate representations of the original documents submitted by the applicant.

Defects in the images include but are not limited to the items checked:

- ☐ BLACK BORDERS
- ☐ IMAGE CUT OFF AT TOP, BOTTOM OR SIDES
- ☒ FADED TEXT OR DRAWING
- ☒ BLURRED OR ILLEGIBLE TEXT OR DRAWING
- ☐ SKEWED/SLANTED IMAGES
- ☐ COLOR OR BLACK AND WHITE PHOTOGRAPHS
- ☐ GRAY SCALE DOCUMENTS
- ☒ LINES OR MARKS ON ORIGINAL DOCUMENT
- ☐ REFERENCE(S) OR EXHIBIT(S) SUBMITTED ARE POOR QUALITY
- ☐ OTHER: \_\_\_\_\_

**IMAGES ARE BEST AVAILABLE COPY.**

**As rescanning these documents will not correct the image problems checked, please do not report these problems to the IFW Image Problem Mailbox.**